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**Measurement of Viscosity
of Gaseous Mixtures at
Atmospheric Pressure**

Jag J. Singh

*Langley Research Center
Hampton, Virginia*

Gerald H. Mall

*Computer Sciences Corporation
Hampton, Virginia*

Hoshang Chegini

*Old Dominion University
Norfolk, Virginia*



National Aeronautics
and Space Administration

Scientific and Technical
Information Branch

Summary

Coefficients of viscosity of various types of gas mixtures, including simulated natural-gas samples, have been measured at atmospheric pressure and room temperature using a modified capillary tube method. Pressure differences across the "straight" capillary tube section of a thermal mass flowmeter were measured for small, well-defined, volume flow rates for the test gases and for standard air. In this configuration, the flowmeter provides the volumetric flow rates as well as a well-characterized capillary section for differential pressure measurements across it. The coefficients of viscosity of the test gases were calculated using the reported value of $185.6 \mu\text{P}$ for the viscosity of dry air at 300 K. The coefficients of viscosity for the test mixtures were also calculated using Wilke's approximation of the Chapman-Enskog (C-E) theory. The experimental and calculated values for binary mixtures are in agreement within the reported accuracy of Wilke's approximation of the C-E theory. However, the agreement for multicomponent mixtures is less satisfactory, presumably because of the limitations of Wilke's approximation of the classical dilute-gas state model.

Introduction

Measurement of viscosity of pure gases and gas mixtures at atmospheric pressure is of great interest in fluid flow and heat transfer studies. Such information is not always readily available for the gas mixtures, particularly for hydrocarbon mixtures. In recent studies (ref. 1) of heats of combustion of natural gas and mixtures of gaseous hydrocarbons, reliable information about their viscosities was needed. In the absence of reliable and accurate experimental data and because of inadequate theoretical understanding of the coefficients of viscosity for these complex mixtures, it was decided to develop experimental procedures for measuring their coefficients of viscosity. It was desired that the technique should be simple and accurate for direct measurement of coefficients of viscosity for all types of gases. It should require only small quantities of the test gas and should be easily adaptable for relative measurements.

There are several techniques for measuring gaseous viscosity. The prominent ones are summarized in the sections which follow.

General Capillary Method

In the case of a compressible fluid flowing steadily through a perfectly cylindrical capillary tube, the volumetric flow rate at the inlet, is given by the

following expression (refs. 2 to 4):

$$V = \frac{\pi r_c^4 (p_i - p_o)}{8\eta\ell_c} \quad (1)$$

where

r_c	root-mean-square radius of tube
ℓ_c	length of tube
η	coefficient of viscosity of fluid
p_i	pressure at inlet
p_o	pressure at outlet

This is Poiseuille's formula for the flow of a gas through a tube. (See ref. 5.) Since all the quantities in the expression are easily measurable, it provides a convenient method for measuring η experimentally. Equation (1) can be simplified as follows:

$$V = A \frac{\Delta p}{\eta} \quad (2)$$

where

$$A = \pi r_c^4 / 8\ell_c$$

A more accurate form of this equation would be (refs. 6 and 7):

$$V = \frac{A \Delta p}{\eta(1 + \delta)} \left(1 + \frac{4\xi}{b}\right) - B \quad (3)$$

where

δ	capillary bore nonuniformity correction factor
$\frac{4\xi}{b}$	wall-slip correction factor, X
B	flow nonlaminarity correction factor

For very low flow rates (corresponding to Reynolds numbers from 100 to 2000) sustained turbulence cannot be expected. Under these conditions, B is almost zero, and the slip correction is negligible for all gases. Hence, equation (2) provides an excellent basis for direct determination of viscosity as follows:

$$\left(\frac{V(\text{Air})}{V(\text{Test})} \right) = \left(\frac{\Delta p(\text{Air})}{\Delta p(\text{Test})} \right) \left(\frac{\eta(\text{Test})}{\eta(\text{Air})} \right) \quad (4)$$

$$\eta(\text{Test}) = \eta(\text{Air}) \left(\frac{V(\text{Air})}{V(\text{Test})} \right) \left(\frac{\Delta p(\text{Test})}{\Delta p(\text{Air})} \right) \quad (5)$$

By measuring the differential pressure across the same capillary tube for low-volume flow rates ($< 50 \text{ cm}^3/\text{min}$) for air and the test gas, the viscosity of the test gas can easily be determined. The

accuracy of the viscosity value is limited only by the accuracies with which V (Air), V (Test), and Δp can be measured.

Oscillating-Disc Method

This method involves measurement of the period and amplitude of the damped oscillations of a suitable solid body. The solid body is suspended from an elastic wire first in the test fluid and then in vacuum. A simplified expression for the fluid viscosity may be written as follows (refs. 8 to 10):

$$\eta = C \left(\frac{\lambda}{T} - \frac{\lambda_o}{T_o} \right) \quad (6)$$

where

λ, λ_o = damping coefficients in test fluid and vacuum, respectively

T, T_o = period of oscillation in test fluid and vacuum, respectively

$C = \frac{4I_d}{\pi r_d^4} \left(\frac{d_1 d_2}{d_i + d_2} \right)$, an experimental constant

I_d = moment of inertia of oscillating disc

r_d = radius of oscillating disc

d_1, d_2 = distances between oscillating and fixed discs

However, the velocity field around the oscillating body is quite complex. This has been the major limitation of this method until recently.

Rotating-Cylinder Method

In its most widely used variant, the angular deflection θ of the inner cylinder is noted when the outer cylinder is rotated with a constant angular velocity ω . If the end effects can be ignored, the coefficient of the fluid viscosity is given by (see refs. 11 and 12):

$$\eta = \frac{\pi \theta I_c (r_o^2 - r_i^2)}{r_o^2 r_i^2 T^2 \omega \ell_i} \quad (7)$$

where

I_c = moment of inertia of inner cylinder

ℓ_i = length of inner cylinder

T = natural period of vibration

r_i, r_o = radii of inner and outer cylinders, respectively

The major limitations of this method are the complexity of calculation of end-effect corrections and the

requirement that no eddy motion develop in the fluid flow between the two cylinders.

The Falling-Body Method

The basis of the method lies in Stokes' law (ref. 13), which states that the viscous drag W for a sphere falling with a uniform velocity in a fluid is given by

$$W = 6\pi\eta r_s V_T \quad (8)$$

where

r_s = radius of falling sphere

V_T = terminal uniform velocity

By measuring the terminal velocity of a sphere of known dimensions, the viscosity of the test medium can be obtained from the following expression:

$$\eta = \frac{2}{9} \left(\frac{\rho_s - \rho_f}{V_T} \right) g r_s^2 \quad (9)$$

where

ρ_s = density of falling sphere

ρ_f = density of test fluid

g = acceleration due to gravity

The major limitations of this method include the problems of frictional charge and the establishment of a truly uniform fall through the test fluid.

From these various techniques, it was decided to use a variant of the capillary flow method for measuring the viscosities of the test gases. This method is particularly suited for complex mixtures, which can often be costly and are available only in small quantities.

As indicated previously, the capillary tube viscometer is based on the application of Poiseuill's law for the steady flow of Newtonian fluids through a capillary tube. Clearly, all the assumptions made in the derivation of the laminar-flow solution cannot be fulfilled in its practical design and operation. However, a detailed study of all the sources of errors indicates that a practical capillary tube viscometer, capable of an accuracy of the order of 1 percent, can be developed if the following conditions are met:

1. A laminar steady-state flow condition is achieved before the differential pressure and volumetric flow rates are measured.

2. The fluid flow rates are kept low (i.e., Reynolds number is kept below 2×10^3).

3. The effects of compressibility of actual fluids flowing through the capillary tube are minimized.

(This is possible if the flow rates are kept below $50 \text{ cm}^3/\text{min}$ and if $\frac{p_o}{p_i} \approx 1$.)

The experimental procedures and the results for a number of binary and multicomponent mixtures are discussed in the following sections.

Symbols

B	flow nonlaminarity constant
d	distance between oscillating disc and fixed disc
f_i	mole fraction of i th specie
I_c	moment of inertia of inner cylinder
I_d	moment of inertia of oscillating disc
ℓ_c	length of capillary tube
ℓ_i	length of inner cylinder
M_i	molecular weight of i th specie
n	number of components in multicomponent mixture
p	pressure
p_i	pressure at inlet of capillary tube
p_o	pressure at outlet of capillary tube
r_c	root-mean-square radius of capillary tube
r_d	radius of oscillating disc
r_i	radius of inner cylinder
r_o	radius of outer cylinder
r_s	radius of falling sphere
T	period of oscillation in test fluid
T_o	period of oscillation in vacuum
V	volumetric flow rate
V_T	terminal uniform velocity
W	viscous drag
X	wall-slip correction factor
Δp	differential pressure ($p_i - p_o$)
Δp_C	pressure drop across series capillary tube
Δp_F	pressure drop across capillary section of flowmeter
δ	capillary bore nonuniformity correction factor

η	coefficient of viscosity
η_i	coefficient of viscosity of i th component
θ	angular deflection of inner cylinder
λ	damping coefficient in test fluid
λ_o	damping coefficient in vacuum
ρ_f	density of test fluid
ρ_s	density of falling sphere
ϕ_{ij}	Chapman-Enskog parameter for species i and j
ω	angular velocity

Theory

The theory of viscosity of gases has been discussed extensively by a number of authors. (See refs. 14 to 17.) In this report, the computational procedure used for calculating the viscosity of multicomponent gas mixtures is briefly reviewed.

According to the Chapman-Enskog (C-E) kinetic theory of gases (ref. 15), the viscosity of a low-pressure multicomponent mixture can be approximated by the following expression:

$$\eta(\text{Mixture}) = \sum_{i=1}^n \frac{\eta_i f_i}{\sum_{j=1}^n f_j \phi_{ij}} \quad (10)$$

where

n	number of components
f_i	mole fraction of i th specie

There are a number of methods for estimating the parameter ϕ_{ij} . We have used Wilke's approximation (refs. 18 to 20) for ϕ_{ij} as follows:

$$\phi_{ij} = \frac{\left[1 + \left(\frac{\eta_i}{\eta_j} \right)^{1/2} \left(\frac{M_j}{M_i} \right)^{1/4} \right]^2}{\left[8 \left(1 + \frac{M_i}{M_j} \right) \right]^{1/2}} \quad (11)$$

$$\phi_{ji} = \left(\frac{\eta_j}{\eta_i} \frac{M_i}{M_j} \right) \phi_{ij} \quad (12)$$

where

f_i	mole fraction of i th component
M_i	molecular weight of i th component

For nonpolar mixtures, Wilke's approximation is preferable over Brokaw's treatment (ref. 21) because

of its simplicity. It has been reported to be accurate to within 1 percent for nonpolar binary mixtures. For example, for a binary mixture,¹ one obtains the following:

$$\eta(\text{Mixture}) = \frac{\eta_1 f_1}{f_1 + f_2 \phi_{12}} + \frac{\eta_2 f_2}{f_2 + f_1 \phi_{21}} \quad (13)$$

It is apparent that the viscosity of the mixture is not equal to the weighted mean of the values of the viscosity of its components. Equation (10) has been used for calculating $\eta(\text{Mixture})$ for various types of mixtures. The viscosities of pure gases have been taken from references 6 and 23. The calculational program is described in the next section.

Computational Procedure

The computer program WILKE is written in FORTRAN Version 5 language for the Control Data CYBER 170 series digital computer system with Network Operating System (NOS) 2.3. The program requires approximately 26 000 octal locations of core storage. A typical run to analyze 30 cases, each with 5 component gases, requires approximately 1 CPU second on the CYBER 173.

Program execution is performed by case, and the user may provide data for as many cases as desired. For each case, data are input to the program in NAMELIST format using the name VAR. Four variables must be defined for each case. The integer case number (NCASE) is used only for bookkeeping. The integer number of gases to be considered (NGAS) may range from 2 to 12. An integer array (IPOINT) is used for selecting which gases are to be included. A value must be provided for each of the NGAS gases. Gases are selected by specifying the appropriate indices from table A. CONC is a real array defining the concentration of each gas. The sum of all concentration fractions should equal 1, and the position within the array for each concentration should also follow the gas/index values provided above. As an example, a run to analyze two cases is presented below. The first case consists of a mixture of 30% CH₄, 50% N₂, and 20% CO₂. The second case consists of a mixture of 35% CH₄, 40% C₂H₆, and 25% C₄H₁₀.

¹ For multicomponent mixtures, Buddenberg and Wilke (ref. 22) recognized the necessity of introducing the diffusion coefficients into the formula for $\eta(\text{Mixture})$ to take into account the collisions between unlike pairs. Even though this modified expression has been reported to predict experimental results quite well, it is not always usable because of lack of data about diffusion coefficients.

Gas	Index
CH ₄	1
C ₂ H ₆	2
C ₃ H ₈	3
N ₂	4
C ₄ H ₁₀	5
CO ₂	6
O ₂	7
NH ₃	8
Ne	9
Ar	10
H ₂	11
He	12

Table A

```
$VAR NCASE = 1, NGAS = 3, IPOINT = 1, 4, 6, CONC =
  0.30, 0.00, 0.00, 0.50, 0.00, 0.20$
```

```
$VAR NCASE = 2, NGAS = 3, IPOINT = 1, 2, 5, CONC =
  0.35, 0.40, 0.00, 0.00, 0.25$
```

Program output for each case echoes the case number and concentrations of selected component gases and gives the viscosity of the mixture. A listing of the program appears on the following page.

Experimental Procedure and Results

It was the purpose of this study to develop a technique for measuring the viscosity of the unknown test gas without knowing its physical characteristics. This required the use of a dry test meter in series with an appropriate mass flowmeter for determining the calibration constant of the flowmeter for the test gas. Using this calibration constant, the volumetric flow rate of the test gas through any other mass flowmeter based on the same principle can easily be determined. As indicated previously, a variant of the capillary tube method has been selected for measuring the viscosity of various types of gaseous mixtures. This method is simple, sensitive, and adaptable to absolute or relative viscosity measurements for low-pressure gases. Figure 1 is a schematic of the experimental system. The capillary section of the flowmeter is made up of two narrow (0.030-cm and 0.045-cm i.d.) capillary tubes in parallel. The flow through each element of the capillary section remains laminar under low flow rates. Figure 2 shows the capillary section of the flowmeter selected for this study. The pressure difference across the capillary section of the selected flowmeter is measured for air and then

```

PROGRAM WILKE(INPUT,OUTPUT,TAPE5=INPUT,TAPE6=OUTPUT)
DIMENSION MASS(15),CONC(15),ETA(15)
DIMENSION DETA(15)
REAL MASS
REAL K,MU
DIMENSION SYMB(15)
DIMENSION IPOINT(15)
CHARACTER*10 SYMB
NAMELIST/VAR/IPOINT,NGAS,CONC,NCASE
DATA MASS/16.,30.,44.,28.,58.,44.,32.,17.,20.,40.,2.,4.,3*0./
DATA ETA/109.8E-6,91.0E-6,80.0E-6,175.69E-6,74.8E-6,
1 146.63E-6,203.31E-6,97.4E-6,313.81E-6,
2 222.86E-6,88.73E-6,196.14E-6,3*0./
DATA DETA/0.330,0.277,0.220,0.454,0.237,
1 0.450,0.616,0.425,0.697,0.704,0.200,0.464,
2 3*0./
DATA SYMB/'CH4','C2H6','C3H8','N2','C4H10','CO2','O2',
1 'NH3','NE','AR','H2','HE(4)',3*1'/

```

CHANGE THIS CARD FOR A DIFFERENT TEMPERATURE

```

DATA T/300./
1 FORMAT(1H1)
2 FORMAT(///14X,'CASE NUMBER ',I2///)
3 FORMAT(///9X,'COMPONENT',6X,'FRACTION'///)
4 FORMAT(9X,A10,6X,F7.4)
5 FORMAT(//3X,'VISCOSITY OF MIXTURE = ',1PE11.3)
DO 8 IJK=1,15
DETA(IJK) = DETA(IJK)*1.E-6
8 CONTINUE
10 CONTINUE
READ(5,VAR,END=20)
20 IF(EOF(5).NE.0) GO TO 999
SUM = 0.
DO 40 I=1,NGAS
ANS = 0.
DO 30 J=1,NGAS
IF(I.EQ.J) GO TO 30
IND1 = IPOINT(I)
IND2 = IPOINT(J)
ETA1 = ETA(IND1) + DETA(IND1)*(T - 293.15)
ETA2 = ETA(IND2) + DETA(IND2)*(T - 293.15)
PHI = 1. + SQRT(ETA1/ETA2)*(MASS(IND2)/MASS(IND1))
1 **0.25
PHI = PHI*PHI
PHI = PHI/SQRT(8.*(1. + MASS(IND1)/MASS(IND2)))
ANS = ANS + CONC(IND2)*PHI
30 CONTINUE
SUM = SUM + ETA1/(1. + ANS/CONC(IND1))
40 CONTINUE
WRITE(6,1)
WRITE(6,2) NCASE
WRITE(6,3)
DO 50 I=1,NGAS
IND = IPOINT(I)
WRITE(6,4) SYMB(IND),CONC(IND)
50 CONTINUE
WRITE(6,5) SUM
GO TO 10
999 CONTINUE
STOP
END

```

for the test gas.² Care is taken to keep the flow rates for the air and the test gas below 50 cm³/min to ensure laminarity of the gas flows through the capillary section of the flowmeter. Figure 3 shows a correlation between the pressure difference across the test capillary section and the flow rate through it. It is obvious that, as would be expected for laminar flows from equation (2), the relationship between the flow rates and the corresponding differential pressures is linear for flow rates from 10 to 50 cm³/min. In this experimental arrangement, the flowmeter provides data for determining the volume flow rates, as well as a well-characterized capillary tube section for differential pressure measurement across it. It is this dual role of the flowmeter that constitutes the uniqueness of the method developed in this study. The results for binary and more complex mixtures are summarized below. All measurements were made at room temperature and at a pressure of 1 atmosphere. Figure 4 is a photograph of the viscosity measurement station.

Purity of gases. Table B lists the components and their purity as quoted by the suppliers.

Gas	Purity
Air	Dry grade; Dew point = -80°F
Ar	99.998%
CO ₂	99.50%
N ₂	99.998%
CH ₄	99.9995%
C ₂ H ₆	99.00%
C ₃ H ₈	99.95%
C ₄ H ₁₀	99.50%

Table B

Preparation of binary mixtures. The binary mixtures were prepared by mixing appropriate volume flow rates of the individual gases. This procedure ensured complete interdiffusion of the two components. Selected gases were mixed in various ratios to produce a total flow rate from 20 to 50 cm³/min. Differential pressure measurements across the test capillary element of the gas flowmeter were made with a differential capacitance gage. Viscosity values for the test mixtures were then calculated by using equation (5). The results are summarized in table I.

² Pressure differences were measured across flowmeter 1 and across series capillary tube section C-1 (fig. 1). The values of the viscosity of pure gases calculated on the basis of these two independent pressure differences were found to be equal within approximately ± 1 percent. It was therefore decided to eliminate the series capillary section from all further measurements.

Preparation of complex mixtures. The more complex mixtures were prepared by a commercial vendor according to our specifications. The vendor prepared the mixtures by gravimetrically mixing different mass fractions, which were later expressed in appropriate mole fractions. The test mixtures were kept in their containers for several weeks before use to ensure complete mixing.

Four different mixtures containing different volume fractions of C_xH_y hydrocarbons as well as N₂ and CO₂ were prepared. Their viscosities were measured using the same procedure described for binary mixtures in the preceding section. The results are summarized in table II.

Natural gas. A test sample of the natural gas was obtained from the Virginia Natural Gas Company. Its coefficient of viscosity was measured by using the same procedure as for the binary and complex mixtures. The test sample was also analyzed for its composition. The composition data are summarized in table III. These values are averages of two measurements of the same sample. The accuracies of individual gas chromatographic analyses are of the order of 1 percent (relative). Using these data, the coefficient of viscosity of the sample was calculated, and the results are summarized in table IV.

Pure gases. Even though the main purpose of this study was to develop a technique for measuring the viscosity of multicomponent gas mixtures, it was deemed desirable to measure the viscosity of selected pure gases as a test of the accuracy of the proposed technique. These selected pure gases were treated as unknown test gases, and their viscosities were measured by using the same procedure as was used for the multicomponent mixtures. The present results, along with the values reported in the literature are summarized in table V. It is apparent that the two sets of data are in excellent agreement. All measured values are strongly dependent on the viscosity of air at room temperature and pressure. The viscosity of dry air has been taken to be 185.6 μ P. (See refs. 6 and 23.)

Discussion

As seen from the data in table I, the agreement between the measured and calculated values of viscosity for binary mixtures is quite good. Also included in the table are some of the values reported by other authors. These values are also in good agreement with experimental values. Table II summarizes the experimental and calculated values for four complex mixtures prepared to simulate U.S. natural-gas samples. In these cases, the agreement is less

satisfactory—possibly because of uncertainties in the mixture compositions. To further resolve this issue, the viscosity of a Virginia Natural Gas sample was measured. This sample was also analyzed for its composition. The measured value of its viscosity using the present technique (independent of gas composition) is $106.7 \mu P$ (± 2.0): the expected value was $107.9 \mu P$ (± 2.0). These values should be compared with a calculated value of $111.6 \mu P$ (± 2.0) based on the listed gas composition. These results indicate that η (calculated) values for complex mixtures containing hydrocarbons are slightly different from η (Measured) values, possibly because of the limitations of Wilke's approximation of the classical dilute-gas state model. A direct measurement of the viscosity of the unknown test mixture may therefore be preferable to the one computed on the basis of its measured composition, since the limitations of Wilke's approximation of the dilute-gas state model make the latter less accurate for complex gas mixtures.

Concluding Remarks

A variant of the general capillary method for measuring the viscosity of unknown gases has been developed. It is based on the use of a thermal-mass flowmeter whose critical sensing element simultaneously provides a capillary section for direct measurement of pressure drop. It is this dual role of the thermal-mass flowmeter that constitutes the uniqueness of the new technique. It does not require any prior information about the composition or other physical properties of the test gases if a dry test meter is used for initial calibration constant determination. It is particularly suited to very complex hydrocarbon mixtures where classical theory limitations as well as compositional errors make theoretical calculations less reliable. The results for pure gases indicate that the technique is accurate to within ± 1 percent. In fact, the only limiting factors appear to be the accuracies with which volume flow rates and differential pressure can be measured.

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Hampton, VA 23665-5225
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References

1. Singh, Jag J.; Sprinkle, Danny R.; and Puster, Richard L.: *New Method for Determining Heats of Combustion of Gaseous Hydrocarbons*. NASA TP-2531, 1985.
2. Flynn, G. P.; Hanks, R. V.; LeMaire, N. A.; and Ross, J.: Viscosity of Nitrogen, Helium, Neon, and Argon From -78.5° to 100°C Below 200 Atmospheres. *J. Chem. Phys.*, vol. 38, no. 1, Jan. 1, 1963, pp. 154-162.
3. Fryer, G. M.: A Theory of Gas Flow Through Capillary Tubes. *R. Soc. London Proc., ser. A.*, vol. 293, July-Aug. 1966, pp. 329-341.
4. Richardson, H. P.; Cummins, D.; and Guereca, R. A.: Absolute Viscosity Determinations by Means of a Coiled-Capillary Viscosimeter. Data for Helium, Carbon Dioxide Mixtures. *Proceedings of the Fourth Symposium on Thermophysical Properties*, J. R. Moszynski, ed., American Soc. Mech. Eng., c.1968, pp. 372-379.
5. Jeans, James: *An Introduction to the Kinetic Theory of Gases*. Cambridge University Press, 1967, pp. 167-170.
6. Touloukian, Y. S.; Saxena, S. C.; and Hestermann, P.: Viscosity. *Thermophysical Properties of Matter*, Volume 11, IFI/Plenum, c.1975.
7. Carr, N. L.: *Viscosities of Natural Gas Components and Mixtures*. Bull. 23, Inst. Gas Technol. (Chicago, Illinois), 1953.
8. Rietveld, A. O.; and Van Itterbeek, A.: Measurements of the Viscosity of Neon-Argon Mixtures Between 300 and 70°K . *Physica*, vol. 22, 1956, pp. 785-790.
9. Kestin, J.; and Wang, H. E.: Corrections for the Oscillating-Disk Viscometer. *J. Appl. Mech.*, vol. 24, 1957, pp. 197-206.
10. Kestin, J.; and Leidenfrost, W.: The Effect of Moderate Pressures on the Viscosity of Five Gases. *Thermodynamic and Transport Properties of Gases, Liquids and Solids*, McGraw-Hill Book Co., Inc., c.1959, pp. 321-338.
11. Van Wazer, J. R.; Lyons, J. W.; Kim, K. Y.; and Colwell, R. E.: *Viscosity and Flow Measurement*. Interscience Publ., 1963.
12. Carmichael, L. T.; Berry, Virginia; and Sage, B. H.: Viscosity of Hydrocarbons. Methane. *J. Chem. & Eng. Data*, pt. I, vol. 10, no. 1, Jan. 1965, pp. 57-61.
13. Condon, E. U.; and Odishaw, Hugh, eds.: *Handbook of Physics*, Second ed. McGraw-Hill Book Co., c.1967.
14. Hirschfelder, Joseph O.; Curtiss, Charles F.; and Bird, R. Bryon: *Molecular Theory of Gases and Liquids, Corrected Printing With Notes Added*. John Wiley & Sons, Inc., 1967.
15. Reid, Robert C.; Prausnitz, John M.; and Sherwood, Thomas K.: *The Properties of Gases and Liquids*, Third ed. McGraw-Hill Book Co., c.1977.
16. Chapman, Sydney; and Cowling, T. G.: *The Mathematical Theory of Non-Uniform Gases*. Cambridge Univ. Press, 1964.
17. Wilke, C. R.: A Viscosity Equation for Gas Mixtures. *J. Chem. Phys.*, vol. 18, no. 4, Apr. 1950, pp. 517-519.
18. Brokaw, Richard S.: *Viscosity of Gas Mixtures*. NASA TN D-4496, 1968.
19. Buddenberg, J. W.; and Wilke, C. R.: Calculation of Gas Mixture Viscosities. *Ind. Eng. Chem.*, vol. 41, no. 7, July 1949, pp. 1345-1347.
20. Gray, Dwight E., ed.: *American Institute of Physics Handbook*, Third ed. McGraw-Hill Book Co., c.1972.
21. Singh, Jag J.; and Puster, Richard L.: *New Technique for Calibrating Hydrocarbon Gas Flowmeters*. NASA TM-85792, 1984.

22. Kestin, J.; Kobayashi, Y.; and Wood, R. T.: The Viscosity of Four Binary, Gaseous Mixtures at 20° and 30°C. *Physica*, vol. 32, no. 6, June 1966, pp. 1065-1089.
23. DiPippo, R.; Kestin, J.; and Oguchi, K.: Viscosity of Three Binary Gaseous Mixtures. *J. Chem. Phys.*, vol. 46, no. 12, June 15, 1967, pp. 4758-4764.
24. Kestin, J.; and Yata, J.: Viscosity and Diffusion Coefficient of Six Binary Mixtures. *J. Chem. Phys.*, vol. 49, no. 11, Dec. 1, 1968, pp. 4780-4791.
25. Giddings, John G.; Kao, James T. F.; and Kobayashi, Riki: Development of a High-Pressure Capillary-Tube Viscometer and Its Application to Methane, Propane, and Their Mixtures in the Gaseous and Liquid Regions. *J. Chem. Phys.*, vol. 45, no. 2, July 15, 1966, pp. 578-586.

TABLE I. SUMMARY OF COEFFICIENTS OF VISCOSITY FOR BINARY MIXTURES
 $[T = 300 \text{ K}; p = 1 \text{ atm}]$

Gas mixture, percent volume	Volume flow rate, cm ³ /min ($\pm 1\%$)	Differential pressure, torr ($\pm 0.25\%$)	Coefficient of viscosity, μP		
			η (Literature)	$\bar{\eta}$ (Measured) ^a	η (Calculated) ^b
Air	48.9	0.693	185.6		
24.9% N ₂ + 75.1% CO ₂	46.6	.555	^c 158.2 \pm 1.6	156.0	155.6
50.2% N ₂ + 49.8% CO ₂	45.8	.573	^c 165.7 \pm 1.7	163.8	162.3
75.0% N ₂ + 25.0% CO ₂	46.1	.597	^c 172.9 \pm 1.7	169.6	169.8
Air	48.9	0.708	185.6		
25.1% N ₂ + 74.9% Ar	39.6	.683	^d 215.8 \pm 0.2	221.1	216.9
49.8% N ₂ + 50.2% Ar	39.0	.623	^d 204.3 \pm 0.2	204.9	204.9
75.3% N ₂ + 24.7% Ar	38.3	.560	^d 191.5 \pm 0.2	187.3	192.1
87.7% N ₂ + 12.3% Ar	35.9	.520	^d 186.1 \pm 0.2	185.5	185.2
Air	45.9	0.689	185.6		
24.4% Ar + 75.6% CO ₂	51.4	.690	^e 167.5 \pm 1.7	166.1	164.8
48.8% Ar + 51.2% CO ₂	51.0	.762	^e 186.0 \pm 1.9	184.9	181.9
74.1% Ar + 25.9% CO ₂	50.6	.848	^e 205.5 \pm 2.1	207.5	202.8
Air	48.8	0.671	185.6		
25.9% N ₂ + 74.1% CH ₄	44.4	.435		132.2	130.8
50.8% N ₂ + 49.2% CH ₄	44.1	.486		148.8	147.8
75.1% N ₂ + 24.9% CH ₄	45.6	.554		164.0	162.9
Air	49.6	0.708	185.6		
25.9% N ₂ + 74.1% C ₂ H ₆	45.2	.392		112.1	109.4
50.1% N ₂ + 49.9% C ₂ H ₆	45.7	.448		126.7	127.4
75.4% N ₂ + 24.6% C ₂ H ₆	49.3	.566		148.5	149.9
Air	49.6	0.708	185.6		
25.0% CO ₂ + 75.0% CH ₄	44.2	.435	^e 126.5 \pm 1.3	128.0	127.6
50.0% CO ₂ + 50.0% CH ₄	44.6	.473	^e 138.1 \pm 1.4	137.9	138.0
75.0% CO ₂ + 25.0% CH ₄	46.1	.515	^e 145.4 \pm 1.5	145.3	145.1
Air	49.6	0.708	185.6	109.3	
25.0% CO ₂ + 75.0% C ₂ H ₆	43.9	.369			108.1
50.0% CO ₂ + 50.0% C ₂ H ₆	45.4	.436		124.9	121.9
75.0% CO ₂ + 25.0% C ₂ H ₆	48.4	.503		135.1	135.9
Air	49.6	0.708	185.6		
24.8% CO ₂ + 75.2% C ₃ H ₈	36.2	.276		98.7	95.0
49.9% CO ₂ + 50.1% C ₃ H ₈	43.9	.387		113.9	109.9
75.0% CO ₂ + 25.0% C ₃ H ₈	45.5	.462		131.3	127.9

^aSee footnotes on page 10.

TABLE I. Concluded
 $[T = 300 \text{ K}; p = 1 \text{ atm}]$

Gas mixture, percent volume	Volume flow rate, cm^3/min ($\pm 1\%$)	Differential pressure, torr ($\pm 0.25\%$)	Coefficient of viscosity, μP		
			η (Literature)	$\bar{\eta}$ (Measured) ^a	η (Calculated) ^b
Air	48.9	0.702	185.6		
25.0% CH_4 + 75.0% C_2H_6	39.1	.298		98.5	98.0
50.0% CH_4 + 50.0% C_2H_6	38.9	.307		102.1	102.0
74.9% CH_4 + 25.1% C_2H_6	40.2	.330		106.1	106.6
Air	48.9	0.702	185.6		
25.8% CH_4 + 74.2% C_3H_8	33.8	.224	^f 86.8 ^g 86.0	85.7	85.1
50.0% CH_4 + 50.0% C_3H_8	39.8	.285	^f 92.9 ^g 93.0	92.7	93.5
74.9% CH_4 + 25.1% C_3H_8	39.2	.302	^f 100.7 ^g 102.0	99.5	101.4
Air	48.9	0.702	185.6		
25.0% C_2H_6 + 75.0% C_3H_8	24.7	.165		86.4	85.0
51.0% C_2H_6 + 49.0% C_3H_8	36.8	.254		89.2	88.0
74.9% C_2H_6 + 25.1% C_3H_8	39.8	.283		91.9	91.0

^aThe uncertainties in η (Measured) values, determined by the accuracies of the gas flow rate and differential pressure measurements, are within the range of ± 2.5 percent in the present study.

^bThe uncertainties in η (Calculated) values, stemming from uncertainties in the mixture composition, are within the range of ± 2 percent.

^cThese values are the results of measurements of absolute viscosity using an oscillating disc viscometer (ref. 22).

^dThese values are the results of relative measurements of viscosity using an oscillating disc viscometer (ref. 23).

^eThese binary mixture viscosity measurements were made using an oscillating disc viscometer (ref. 24).

^fThese data are cited in reference 15 for comparison between the experimental data and the calculated values. They are in agreement with Wilke's approximation to an accuracy of ± 1 percent.

^gThese are absolute values obtained using a high-pressure capillary tube viscometer. The reproducibility of the data has been reported at ± 0.25 percent (ref. 25).

TABLE II. SUMMARY OF COEFFICIENTS OF VISCOSITY FOR COMPLEX MIXTURES

$$\left[\begin{array}{l} \eta(\text{Air}) = 185.6 \text{ } \mu\text{P}; \\ T = 300 \text{ } K; \\ p = 1 \text{ atm} \end{array} \right]$$

Gas mixture, percent volume ($\pm 1\%$ relative)	Volume flow rate, cm^3/min	Differential pressure, torr ($\pm 0.25\%$)	Coefficient of viscosity, μP	
			$\bar{\eta}(\text{Measured})^a$	$\eta(\text{Calculated})^b$
Air 44.6% C_2H_6 + 25.9% C_3H_8 + 10.8% C_4H_{10} + 5.1% CO_2 + 13.6% N_2	48.9 \pm 0.5	0.682		
	31.8 \pm 1.0	.220	92.0 \pm 1.5	97.0 \pm 2.7
	39.3 \pm 1.2	.318	105.3 \pm 1.1	100.7 \pm 2.8
	36.8 \pm 1.4	.298	105.7 \pm 0.7	102.1 \pm 2.9
51.5% CH_4 + 21.2% C_2H_6 + 23.8% C_3H_8 + 1.7% CO_2 + 1.7% N_2	44.6 \pm 0.9	.308	91.7 \pm 0.7	100.0 \pm 2.8

^aThe errors cited on $\eta(\text{Measured})$ are standard deviations on 5 measurements for each mixture. They do not include the effects of uncertainties in flow rates and differential pressure measurements.

^bThe uncertainties cited on $\eta(\text{Calculated})$ represent the effects of uncertainties in the mixture composition.

TABLE III. COMPOSITION OF NATURAL GAS SAMPLE

[Ref. 21]

Component gas	Volume percentage
CH ₄	96.52 ± 0.20
C ₂ H ₆	2.32 ± 0.08
C ₃ H ₈	0.40 ± 0.03
C ₄ H ₁₀	0.14 ± 0.02
C ₅ H ₁₂	< 0.05
C ₆ H ₁₄	≤ 0.05
CO ₂	0.22 ± 0.02
N ₂	0.39 ± 0.06

TABLE IV. SUMMARY OF COEFFICIENTS OF VISCOSITY FOR NATURAL GAS

η (Measured), μP	η (Calculated), ^a μP
Based on gas composition summarized in table III ^b	Independent of gas composition ^c
107.9 ± 2.0	106.7 ± 2.0
	111.6 ± 2.0

^aThe uncertainty cited on η (Calculated) represents the effects of uncertainties in the mixture composition.

^bThe value measured on the basis of known gas composition is based on the flow-rate calibration constant calculated for that composition. The error cited on it is the standard deviation based on 6 measurements.

^cThe measured value listed in the second column has been determined using the experimentally determined flowmeter calibration constant for the test gas. The error cited on it is the standard deviation based on 6 measurements.

TABLE V. SUMMARY OF COEFFICIENTS OF VISCOSITY FOR PURE GASES

 $[T = 300 \text{ K}; p = 1 \text{ atm}]$

Gas	Volume flow rate cm ³ /min, ($\pm 1\%$)	Differential pressure torr ($\pm 0.25\%$)	Coefficient of viscosity, μP	
			Measured, ^a (average)	Reported ^b (refs. 6 and 23) ($\pm 2\%$)
Air	48.9	0.708		185.6
N ₂	32.3	.449	178.1 \pm 0.5	178.6
CO ₂	35.6	.412	150.3 \pm 0.1	149.9
Ar	41.0	.723	228.2 \pm 0.2	227.6
CH ₄	30.9	.260	112.7 \pm 2.2	112.0
C ₂ H ₆	37.1	.266	95.7 \pm 2.2	94.5
C ₃ H ₈	30.1	.198	83.9 \pm 1.6	82.5
C ₄ H ₁₀ (iso)	23.5	.141	76.6 \pm 1.0	75.7

^aThe errors cited on the measured values are standard deviations based on 3-5 measurements.

^bThe uncertainties on the reported values fall within the range of ± 2 percent.

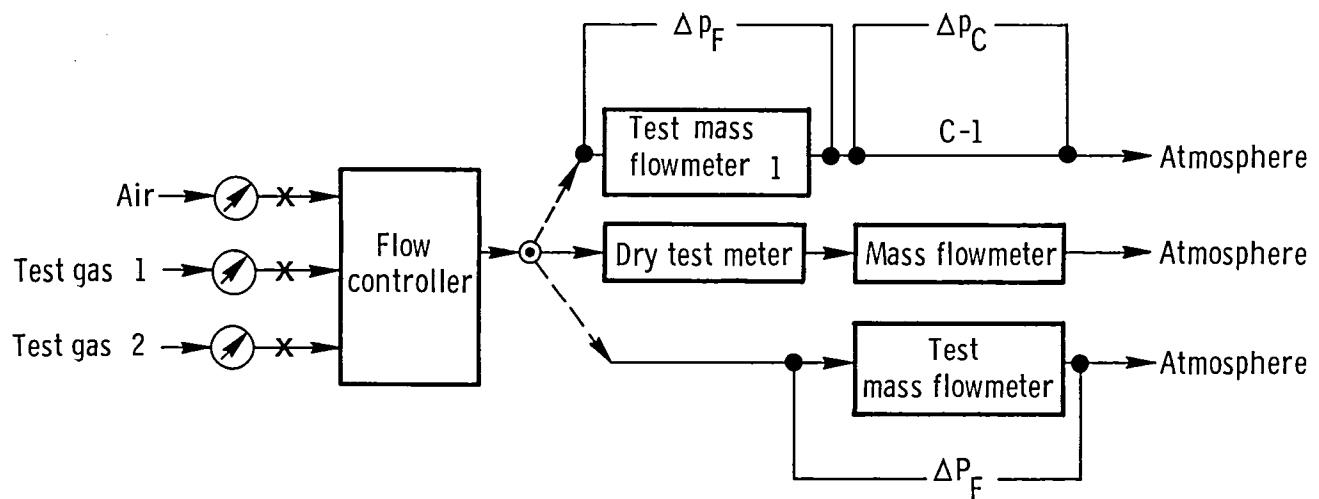


Figure 1. Schematic of experimental system for measuring viscosities of atmospheric-pressure gases.

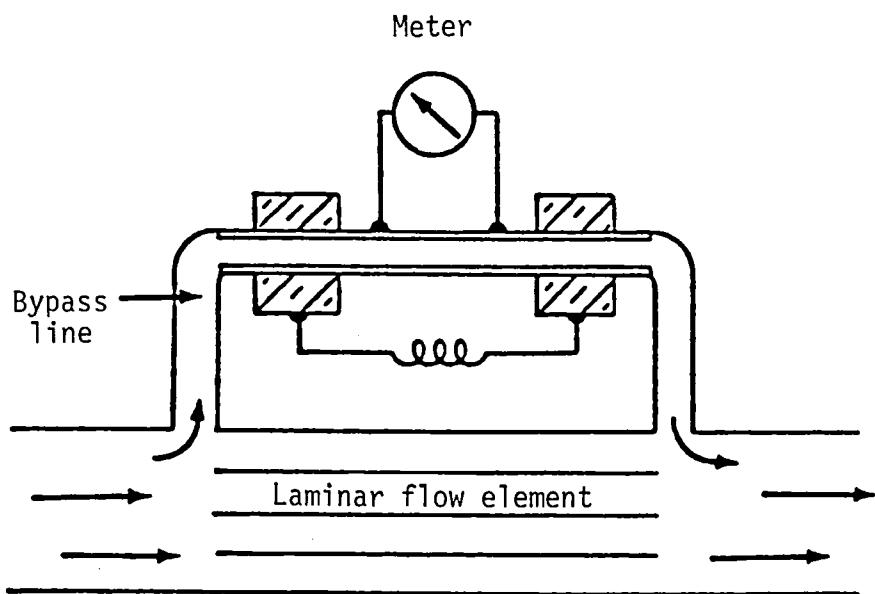


Figure 2. Thermal flowmeter with laminar flow element.

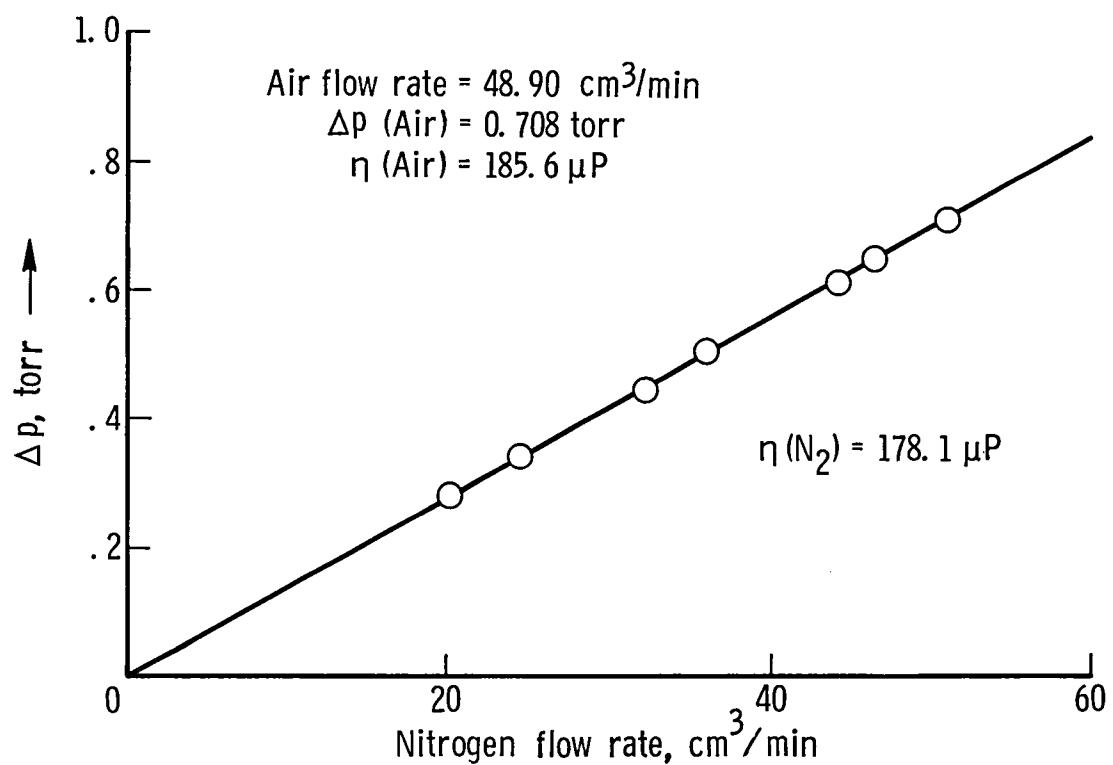
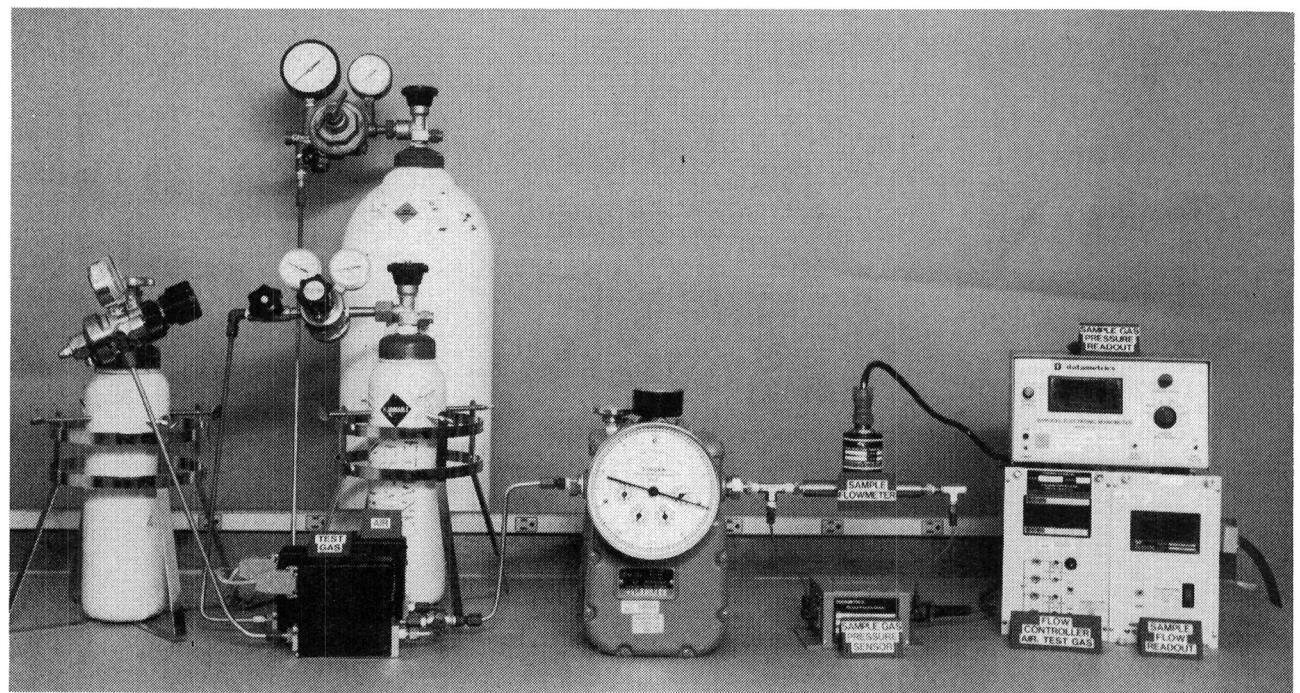


Figure 3. Flow rate versus differential pressure across test capillary section of flowmeter.



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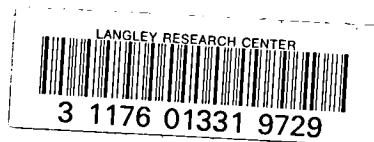
Figure 4. Photograph of experimental viscosity measurement station.

Standard Bibliographic Page

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16. Abstract Coefficients of viscosity of various types of gas mixtures, including simulated natural-gas samples, have been measured at atmospheric pressure and room temperature using a modified capillary tube method. Pressure drops across the "straight" capillary tube section of a thermal mass flowmeter were measured for small, well-defined, volume flow rates for the test gases and for standard air. In this configuration, the flowmeter provides the volumetric flow rates as well as a well-characterized capillary section for differential pressure measurements across it. The coefficients of viscosity of the test gases were calculated using the reported value of $185.6 \mu\text{P}$ for the viscosity of air. The coefficients of viscosity for the test mixtures were also calculated using Wilke's approximation of the Chapman-Enskog (C-E) theory. The experimental and calculated values for binary mixtures are in agreement within the reported accuracy of Wilke's approximation of the C-E theory. However, the agreement for multicomponent mixtures is less satisfactory, possibly because of the limitations of Wilke's approximation of the classical dilute-gas state model.				
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